

# Study Guide

## Science 9 Chemistry TEST

Up until naming (that's on test 2)

Fill in the Bohr Diagrams:

H								He
Li	Be	B	C	N	O	F	Ne	
Na	Mg	Al	Si	P	S	Cl	Ar	
K	Ca							

Be able to define heterogeneous and homogeneous with examples.

What is an element? A chemical?

What can you learn about a product that says 'Chemical Free'?



Where are metals and non-metals located on the periodic table?  
What are some common characteristics of metals?

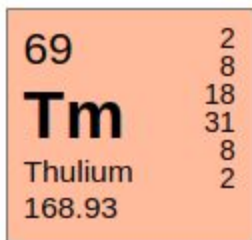
What was so amazing about Dmitri Medeleev's periodic table?  
Why is he considered a genius for constructing this table?

What is a valence shell? Valence electron?  
How can you tell from the location on the periodic table how many valence electrons an element has?  
From the number of valence electrons, what can you infer about how reactive an element is?

List and define the 3 subatomic particles that compose an element.

What is an ion?  
What are the two types?

Looking at the periodic table; what can you tell about the relative sizes of the atoms?  
From left to right, and from top to bottom.  
-- Be able to explain why.



What is the atomic number?  
What is the symbol?  
What is the name?  
What is the atomic mass?  
What do the numbers on the right mean?  
\*This element was chosen for Mr. Harris. He knows why.\*

Complete the following table:

Symbol	Atomic Number	#Protons	#Electrons
Ne			
	22		
		19	

Naming:

Formula	Name	Type (simple ionic, multivalent ionic, polyatomic ion or covalent)
	Magnesium Chloride	
$\text{Li}_2\text{O}$		
	Copper (II) Sulphide	
$\text{MnF}_4$		
	Sodium Hydrogen Carbonate	
$\text{Ni}(\text{NO}_3)_2$		
	Carbon tetrafluoride	
$\text{P}_2\text{O}_5$		